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1 CK-12 Chemistry Concepts - Intermediate Answer Key Chapter 12: Stoichiometry 12.1 Everyday Stoichiometry Practice Questions Use the link below to answer the following questions: 1. What does stoichiometry help you figure out? 2. What are all reactions dependent upon? 3. If I have ten hydrogen molecules and three oxygen molecules, how many molecules of water can I make?

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Stoichiometry 379 CHAPTER 12 Assessment 36 Chapter 12 stoichiometry test b answer key. a. Two formula units  $KClO_3$  decom- pose to form two formula units  $KCl$  and three molecules  $O_2$ . b Chapter 12 stoichiometry test b answer key. Four molecules  $NH_3$  react with six molecules  $NO$  to form five mol-

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Chapter 12 Stoichiometry127 SECTION 12.1 THE ARITHMETIC OF EQUATIONS (pages 353–358) This section explains how to calculate the amount of reactants required or product formed in a nonchemical process. It teaches you how to interpret chemical equations in terms of interacting moles, representative particles, masses, and gas volume at STP.

### SECTION 12.1 THE ARITHMETIC OF EQUATIONS

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Answer:  $4.93 \times 10^{-5}$  L or 49.3  $\mu$ L In Example 12.2.1 and Example 12.2.2, the identity of the limiting reactant has been apparent:  $[Au(CN)_2]^-$ ,  $LaCl_3$ , ethanol, and para -nitrophenol. When the limiting reactant is not apparent, we can determine which reactant is limiting by comparing the molar amounts of the reactants with their ...

### Chapter 12.2: Stoichiometry of Reactions in Solution ...

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