

## Kinetic Molecular Theory Pogil 2005 Answers

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### Kinetic Molecular Theory Pogil 2005

Kinetic Molecular Theory ©POGIL - 2005 5/5 Authored by Applications 1. There is a government warning on all aerosol cans that states: Do not store at a temperature above 120° F (50°C). a) Explain why this warning is required in terms of the relationship between temperature and pressure and the kinetic molecular theory.

### Kinetic Molecular Theory - Ms. Lynch's Lessons

POGIL: Kinetic Molecular Theory. Modified from Foundations of Chemistryby David HansonPage 1 of 3. POGIL: Kinetic Molecular Theory. Learning Objectives. Identify the basic differences between particle behavior in the solid, liquid, and gaseous phases. Develop an understanding of the postulates of the kinetic molecular theory. Identify how temperature affects molecular motion. Apply the kinetic molecular theory to predict the outcome of everyday situations.

### POGIL: Kinetic Molecular Theory

The kinetic molecular theory (KMT) is a simple microscopic model that effectively explains the gas laws described in previous modules of this chapter. This theory is based on the following five postulates described here. (Note: The term "molecule" will be used to refer to the individual chemical species that compose the gas, although some ...

### 9.5 The Kinetic-Molecular Theory - Chemistry 2e | OpenStax

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### Answers Kinetic Molecular Theory Pogil

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### 9.5 The Kinetic-Molecular Theory - Chemistry

Information (Kinetic Molecular Theory) The observed behavior of gases, as expressed by the empirical gas laws, can be understood on the basis of the kinetic molecular theory, which developed over many years up to 1857, when Rudolf Clausius (1822-1888) published it in its most complete and elegant form.

### Chem 116 POGIL Worksheet - Week 2 Gas Laws - Part 2

The Kinetic Molecular Theory and Graham's Laws. The kinetic molecular theory can be used to explain the results Graham obtained when he studied the diffusion and effusion of gases. The key to this explanation is the last postulate of the kinetic theory, which assumes that the temperature of a system is proportional to the average kinetic energy ...

### The Kinetic Molecular Theory - Purdue University

The average kinetic energy of the gas particles is directly proportional to the Kelvin temperature of the gas. 3/3 ©POGIL -2005 Authored by Edited by Linda Padwa and David Hanson, Stony Brook University Kinetic Molecular Theory Key Questions 1. What causes a gas to exert pressure when confined in a container? of 2.

### MRS SCICCHITANO - Home

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### Ms. Demonte's Chemistry Classes - Home

Kinetic Molecular Theory states that gas particles are in constant motion and exhibit perfectly elastic collisions. Kinetic Molecular Theory can be used to explain both Charles' and Boyle's Laws. The average kinetic energy of a collection of gas particles is directly proportional to absolute temperature only.

### Kinetic Molecular Theory | Boundless Chemistry

The temperature of a gas depends on its average kinetic energy  $\text{avg}(1/2mv^2) = 3/2kT$ . In other words, the energy of an ideal gas is entirely kinetic. The amazing thing about the kinetic molecular theory is that it can be used to derive the ideal gas law.

### Kinetic Molecular Theory: The Kinetic Molecular Theory ...

The Kinetic-Molecular Theory {Boyle's Law  $zP \propto 1/V$  zAs the V increases the molecular collisions with container walls decrease and the P decreases. {Dalton's Law  $zP_{\text{total}} = P_A + P_B + P_C + \dots$  zBecause gases have few intermolecular attractions, their pressures are independent of other gases in the container. {Charles' Law  $zV \propto T$

### CHAPTER 12 GASES AND KINETIC-MOLECULAR THEORY

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